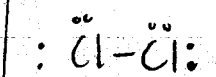
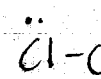
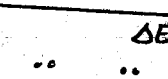
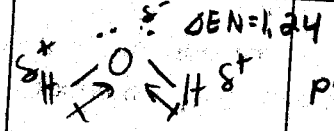
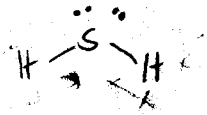
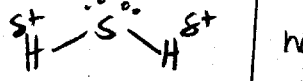
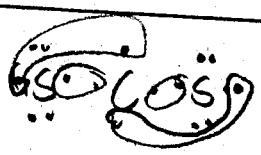
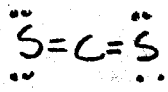
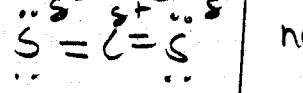
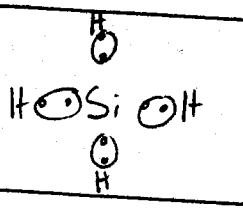
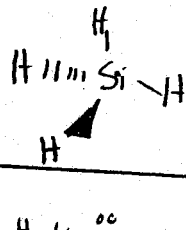
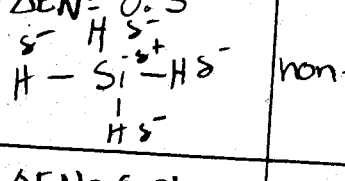
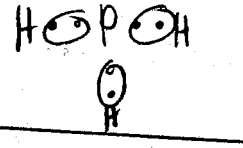
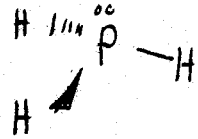
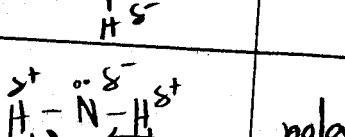
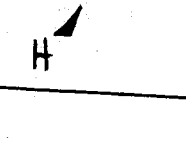
SCH 3UI **Investigation 2-B Modelling Molecules**

Complete the chart below using the compounds listed on the left side.

Compound	Draw the Lewis Structure	3-D -using the model kits create a 3-D version of compound -Draw in below	Name of Shape -Using the 3D molecule write the name of the shape below	Calculate the ΔEN for each type of bond in molecule -draw in partial charges and vector where appropriate	Polar OR Non polar Molecule,
a) H_2			linear	ΔEN=0 	non-polar
b) Cl_2			linear	ΔEN=0 	non-polar
c) H_2O			bent	ΔEN=1.24 	polar
d) H_2S			bent	ΔEN=0.38 	non-polar
e) CS_2			linear	ΔEN=0.03 	non-polar
f) SiH_4			tetrahedral	ΔEN=0.3 	non-polar
g) PH_3			pyramidal	ΔEN=0.01 	non-polar
h) NH_3			pyramidal	 ΔEN=0.84	polar