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Unit Outline:
Quantities in Chemistry

DAY BREAKDOWN

| Date | Pages | Topics | Homework/Assignments | Evaluation |
| :---: | :---: | :---: | :---: | :---: |
| 1 | 76-79 | Getting Started <br> p. 76-77 \# 1-11 <br> - (Ac) Mineral Content of Milk | p. 76-77 \# 1-11 <br> - (Ac) Mineral Content of Milk |  |
| 2 | 80-84 | Topic: Amounts in Chemistry LG: I can calculate the average atomic mass and calculate moles, given mass and molar mass <br> - (N) Mass, Moles, Molar Mass <br> - (Ac) Counting by Mass p. 91 | - (Ac) Counting by Mass <br> p. 91 <br> p. 92 \# 1-4 |  |
| 3 | 85-92 | Topic: Amounts in Chemistry LG: I can calculate molar mass and mass given moles and a chemical formula <br> - (N) Periodic Table and Mass of Elements <br> - (WS) Ratios in Chemical Equations, Amounts in Chemistry | - (WS) Ratios in Chemical Equations, Amounts in Chemistry p. 92 \# 5-10 | (Q) Intro to mass, moles \& molar mass |
| 4 | 93-102 | Topic: Calculations Involving the Mole LG: I can calculate number of atoms and molecules given moles <br> - (N) Calculations Involving Atoms, Molecules \& Molar Mass <br> - PRACTICE p. 94 \# 1-15 | - PRACTICE p. 94 \# 115 |  |
| 5 | 103-105 | Topic: Calculations Involving the Mole LG: I can calculate \# atoms/molecules from moles or mass of substance <br> - (N) Calculating \# of Atoms from Mass - (Ac) p. 105 Counting Atoms, Molecules, and Other Entities | - (Ac) p. 105 Counting Atoms, Molecules, and Other Entities -p. 106 \# 7-9 | (Q) Calcs. Involving the Mole. |
| 6 | 107-120 | Topic: Determining Chemical Formulas LG: I can calculate EF, MF and percent composition <br> - (N) Empirical Formula, Molecular Formula and Percent Composition <br> - (Ac) p. 110 What Makes Popcorn Pop? <br> - PRACTICE p. 114 \# 1-7 | - (Ac) p. 110 What Makes Popcorn Pop? <br> - PRACTICE p. 114 \# 17 |  |
| 7 |  | Topic: Percent Composition by Mass LG: I can calculate EF, MF and percent composition <br> - (L) Percent Composition of Double Bubble | p. 120 \# 1, 3-7 | (Q) Chemical Formulas |
| 8 | 123-130 | Topic: Concentrations of Solutions LG: I can calculate $\mathrm{v} / \mathrm{v}, \mathrm{m} / \mathrm{v}$ and molarity of solutions. <br> - (N) \% Concentration \& Molar Concentrations <br> - PRACTICE p. 126 \# 1-15 | - PRACTICE p. 126 \# 115 |  |

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| 9 | 131-136 | Topic: Concentration of Solutions LG: I can calculate ppm \& $C_{1} V_{1}=C_{2} V_{2}$ <br> - (N) Parts per Million, Dilutions <br> - PRACTICE p. 133 \# 16-22 | $\begin{aligned} & \text { - PRACTICE p. } 133 \text { \# } \\ & 16-22 \\ & \text { p. } 137 \# 2,4 \end{aligned}$ | (Q) [] of Solutions |
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| 12 | 144-148 | Topic: Stoichiometry <br> LG: I can use stoichiometry to predict masses of reactants or products <br> - (N) Stoichiometry <br> - (Ac) Yummy Stoichiometry <br> - (HO) Case Study: Combustion and Carbon Monoxide Poisoning | - (HO) Case Study: <br> Combustion and Carbon Monoxide Poisoning <br> p. 148 \# 1-5 |  |
| 13 | 149-153 | Topic: Limiting \& Excess Reagents LG: I can calculate the limiting and excess reactant. <br> - (N) Limiting \& Excess Reagents <br> - PRACTICE p. 153 \# 1-4 | - PRACTICE p. 153 \# 1- <br> 4 <br> -p. 153 \# 1,2,4,6 |  |
| 14 | 154-155 | Topic: Determining the L.R. and Predicting Mass of Product <br> LG: I can calculate the limiting and excess reactant. <br> - (L) 2.11 The Limiting Reagent and Percentage Composition of a Mixture | - (L) 2.11 The Limiting Reagent and Percentage Composition of a Mixture | - (L) 2.11 |
| 15 | 155-159 | Topic: Percentage Yield LG: I can calculate the percent yield of a reaction. <br> - (N) Percentage Yield <br> - PRACTICE p. 158 \# 1-4 | ```- PRACTICE p. 158 # 1- 4 - p. 159 # 1-4``` |  |
| 16 | 160 | Topic: Calculating the Percent Yield of a Chemical Reaction <br> LG: I can calculate the percent yield of a reaction. <br> - (L) 2.13 Percentage Yield of a Chemical Reaction | - (L) 2.13 | - (L) 2.13 |
| 17 |  | Unit 2 Practice Test |  | Unit 2 Practice Test |
| 18 | 169-173 | Unit 2 Review | $\begin{aligned} & \text {-p. 169-173 \# 1-7, 9, } \\ & 11,12,13-20 \end{aligned}$ |  |
| 19 |  | Unit 2 Test |  | Unit 2 Test |

