BCI SCIENCE SCH 3UI

Unit Outline: Quantities in Chemistry

DAY BREAKDOWN

Date	Pages	Topics	Homework/Assignments	Evaluation	
1	76-79	Getting Started	p. 76-77 # 1-11		
		p. 76-77 # 1-11	- (Ac) Mineral Content		
		- (Ac) Mineral Content of Milk	of Milk		
2	80-84	Topic: Amounts in Chemistry	- (Ac) Counting by Mass		
		LG: I can calculate the average atomic	p. 91		/4
		mass and calculate moles, given mass	p. 92 # 1-4		
		and molar mass			
		- (N) Mass, Moles, Molar Mass			
		- (Ac) Counting by Mass p. 91			
3	85 -92	Topic: Amounts in Chemistry	- (WS) Ratios in	(Q) Intro to mass, moles &	
		LG: I can calculate molar mass and	Chemical Equations,	molar mass	
		mass given moles and a chemical	Amounts in Chemistry		/4
		formula	p. 92 # 5 - 10		
		- (N) Periodic Table and Mass of			
		Elements			
		- (WS) Ratios in Chemical Equations,			
		Amounts in Chemistry			
4	93-102	Topic: Calculations Involving the Mole	- PRACTICE p. 94 # 1-		
		LG: I can calculate number of atoms	15		/4
		and molecules given moles			
		- (N) Calculations Involving Atoms,			
		Molecules & Molar Mass			
		- PRACTICE p. 94 # 1-15			
5	103-105	Topic: Calculations Involving the Mole	- (Ac) p. 105 Counting	(Q) Calcs. Involving the Mole.	
		LG: I can calculate # atoms/molecules	Atoms, Molecules, and		1
		from moles or mass of substance	Other Entities		/4
		- (N) Calculating # of Atoms from Mass	- p. 106 # 7-9		
		- (Ac) p. 105 Counting Atoms, Molecules,			
		and Other Entities			_
6	107-120	Topic: Determining Chemical Formulas	- (Ac) p.110 What		
		LG: I can calculate EF, MF and	Makes Popcorn Pop?		
		percent composition	- PRACTICE p. 114 # 1-		/4
		- (N) Empirical Formula, Molecular	7		
		Formula and Percent Composition			
		- (Ac) p.110 What Makes Popcorn Pop?			
-		- PRACILCE p. 114 # 1-7	400 // 4 0 7		_
/		Topic: Percent Composition by Mass	p. 120 # 1, 3-7	(Q) Chemical Formulas	14
		LG: I can calculate EF, MF and			77
		percent composition			
		- (L) Percent Composition of Double			
0	122 120	BUDDIE			-
ð	123-130	1 opic: Concentrations of Solutions	- PRACILCE p. 126 # 1-		
		LG: I can calculate V/V, m/V and	C1		/4
		molarity of solutions.			
		- (IN) % concentration & Molar			
		DDACTTCE = 126 # 1 15			
		- PRACILCE P. 120 # 1-15			

Date: ____

9	131-136	Topic: Concentration of Solutions LG: I can calculate ppm & C ₁ V ₁ = C ₂ V ₂ - (N) Parts per Million, Dilutions - PRACTICE p. 133 # 16-22	- PRACTICE p. 133 # 16-22 p. 137 # 2, 4	(Q) [] of Solutions	
12	144-148	Topic: Stoichiometry LG: I can use stoichiometry to predict masses of reactants or products - (N) Stoichiometry - (Ac) Yummy Stoichiometry - (HO) Case Study: Combustion and Carbon Monoxide Poisoning	- (HO) Case Study: Combustion and Carbon Monoxide Poisoning p. 148 # 1 - 5		
13	149-153	Topic: Limiting & Excess Reagents LG: I can calculate the limiting and excess reactant. - (N) Limiting & Excess Reagents - PRACTICE p. 153 # 1-4	- PRACTICE p. 153 # 1- 4 - p. 153 # 1,2,4,6		
14	154-155	Topic: Determining the L.R. and Predicting Mass of Product LG: I can calculate the limiting and excess reactant. - (L) 2.11 The Limiting Reagent and Percentage Composition of a Mixture	- (L) 2.11 The Limiting Reagent and Percentage Composition of a Mixture	- (L) 2.11	
15	155-159	Topic: Percentage Yield LG: I can calculate the percent yield of a reaction. - (N) Percentage Yield - PRACTICE p. 158 # 1-4	- PRACTICE p. 158 # 1- 4 - p. 159 # 1-4		
16	160	Topic: Calculating the Percent Yield of a Chemical Reaction LG: I can calculate the percent yield of a reaction. - (L) 2.13 Percentage Yield of a Chemical Reaction	- (L) 2.13	- (L) 2.13	
17		Unit 2 Practice Test		Unit 2 Practice Test	
18	169-173	Unit 2 Review	- p. 169-173 # 1-7, 9, 11, 12, 13-20		
19		Unit 2 Test		Unit 2 Test	