

Investigation 4-A: Making an Activity Series of Metals

Clancey, et. al., (2011) Chemistry 11, McGraw-Hill Ryerson Page 188-189

Title (____/4)

- Original, descriptive title
- name

- partner's name,
- date

/18

Abstract (____/8)

- purpose:
- procedure:
- results:

Comment [MPT3]: Summarize the purpose in ONE sentence

Comment [MPT4]: Summarize the procedure in ONE sentence

Comment [MPT5]: Summarize the results in ONE sentence

Comment [MPT6]: Write a chemical equation that is balanced, and includes states.

Chemical Reactions

- A.
- B.
- C.
- D.
- E.

Observations (____/5)

- Include your observation tables
- Include a written description of your observations

Comment [MPT7]: Students need to record the physical properties, especially colour and state, of the metals & solutions beforehand.

Comment [MPT8]: Remind students of the signs of a chemical reaction include;

- 6) Change in colour
- 7) Formation of a gas
- 8) Formation of a precipitate
- 9) Change in temperature
- 10) Energy produced (e.g. heat, light, sound)

Comment [MPT9]: Write a paragraph describing your observations, which are summarized in your tables. E.g. Observed physical properties of reactants, or observed signs of a chemical reaction

Comment [MPT10]: Write a conclusion based on your purpose

Conclusion: (____/1)

Table 1. Physical Properties

Substance	Observed Physical Properties
EXAMPLE: Strontium (II) nitrate	Transparent (clear), colourless liquid

Table 2. Reactions of metals with various aqueous solutions

Solution (include chemical formula)	Metals			
	Copper (Cu)	Magnesium (Mg)	Iron (Fe)	Zinc (Zn)
EXAMPLE Strontium (II) nitrate	NR	Bubbles formed Solution turned light pink	NR	NR
Copper (II) nitrate				
Magnesium nitrate				
Iron (III) nitrate				
Zinc nitrate				
Hydrochloric acid				
Water H ₂ O				