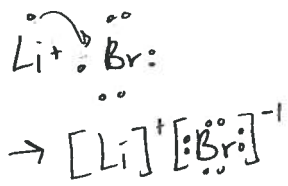
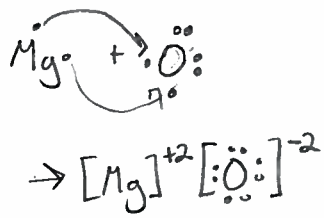
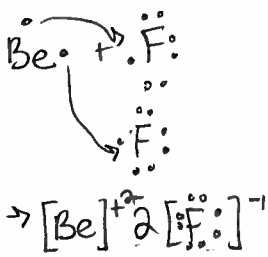
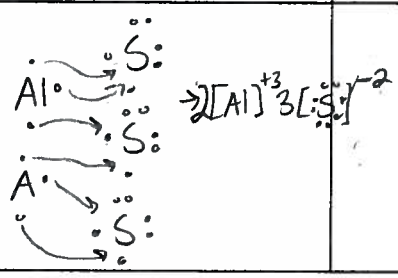
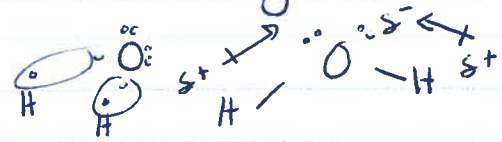
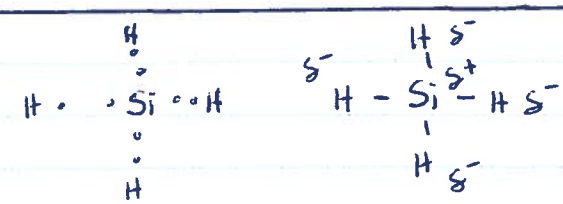
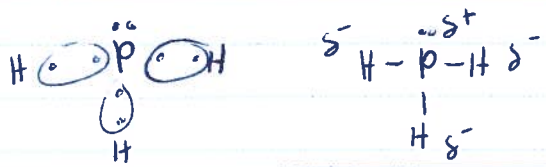
Bonding Atoms	ΔEN	Type of Bond	Formation of Bond (Movement of Electrons)	Ions formed	Chemical Formula
Lithium and Bromine	1.98	Ionic		Li^{+} Br^{-1}	LiBr
Magnesium and Oxygen	2.13	Ionic		Mg^{+2} O^{-2}	MgO
Beryllium and fluorine	2.41	Ionic		Be^{+2} F^{-1}	BeF_2
Aluminium and Sulphur	0.97	Ionic		Al^{+3} S^{-2}	Al_2S_3

Compound	ΔEN	Bond Type	Lewis Diagram / Structural Form
H_2O	1.24	polar covalent	
SiH_4	0.3	non-polar covalent	
PH_3	0.1	non-polar covalent	
~~HO~~	~~0.89~~	~~polar covalent~~	~~  ~~
HCN	$\Delta EN_{CH} = 0.35$ $\Delta EN_{CN} = 0.49$	non-polar covalent	
C_2H_4	$\Delta EN = 0.35$	non-polar covalent